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reaction $\text{CaCO}_3 + 2\text{HNO}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 +$
 $\text{CO}_2 + \text{H}_2\text{O}$, how many milliliters of
0.1015 M HNO_3 are required to react with

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0.9857 grams of CaCO_3 ? (a) 0.1940

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QUESTIONS Select the one best answer
for each question. 1. In the reaction

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$\text{CaCO}_3 + 2\text{HNO}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{CO}_2 + \text{H}_2\text{O}$, how many milliliters of 0.1015 M HNO_3 are required to react with 0.9857 grams of CaCO_3 ? (a) 0.1940 mL (b) 97.04 mL (c) 194.1 mL 2.

**Chemistry 213 Assignment No. 2 Name:
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The following quiz contains 5 multiple choice questions. If you wish to take a longer quiz, please select 'Review Questions' from the navigation bar. ...

Which of the following is not a property required of the substance chosen for use as a precipitate in a gravimetric analysis? A hydrated magnesium salt, $\text{MgSO}_4 \cdot x\text{H}_2\text{O}$

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Questions And Answers For Gravimetric Analysis

CHEM 201 Self Quiz – 2 (Gravimetric

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analysis/Volumetric analysis) Answer key

1. Calcium fluoride, CaF_2 , is a sparingly soluble salt. The concentration of Ca^{2+} ions in a saturated solution is 8.6×10^{-4} M.

$\text{CaF}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{F}^-$ $K_{sp} = [\text{Ca}^{2+}][\text{F}^-]^2$

$[\text{Ca}^{2+}] = [\text{CaF}_2] = 8.6 \times 10^{-4}$

$[\text{F}^-] = 2[\text{CaF}_2] = 2 \times 8.6 \times 10^{-4} = 17.2 \times 10^{-4}$

$K_{sp} = 8.6 \times 10^{-4} \times (17.2 \times 10^{-4})^2 = 1.5 \times 10^{-10}$

Calcium fluoride, CaF_2 , is a sparingly ...

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