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Thermodynamics. Lecture 01.

21. Thermodynamics

~~Thermodynamics, PV Diagrams,
Internal Energy, Heat, Work,
Isothermal, Adiabatic,
Isobaric, Physics 5.1 |~~

~~MSE104 — Thermodynamics of
Solutions~~

Mirza Beg (2019) Day 2.13:

ABC Society, Morality,
Entropy, Phase Rule, Henry
Adams, Social Phase

**Paul
Cockshott - Marx, Babbage
and Boltzmann (SICSA 2011
Keynote) Mod 01 Lec 05**

~~Solution models, chemical
potential~~

Thermodynamics and Heat
transfer Prof S Khandekar

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*and the End of the Universe:
Energy, Entropy, and the
fundamental laws of physics.*

Lec 1 | MIT 5.60

Thermodynamics \u0026

Kinetics, Spring 2008

Second Law of

Thermodynamics, Entropy

\u0026 Gibbs Free Energy FE

Review - Thermodynamics 6.1

| MSE104 - Scheil Equation

Basic Thermodynamics-

Lecture 1 Introduction

\u0026 Basic Concepts MATLAB

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1.1* The plot of $V = V(P, T)$ for a gas is shown in

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Fig. 1.1. Determine the expressions of the two second derivatives of the volume of this plot. (note: the principle curvatures of the surface are proportional to these second derivatives).

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Introduction to the
Thermodynamics of Materials
David R. Gaskell

Preliminaries † Settings
::spellD †
Physical Constants Needed
for Problems ü Heat
Capacities The generic heat
capacity $c_{105} = a +$
 $bT + cT^2 + dT^3 + eT^4 + fT^5 + gT^6 + hT^7 + iT^8 + jT^9 + kT^{10}$

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Manual T2.103 The heat capacities of various elements and compounds are $C_p A_g = C_p \hat{e}$.
8a \emptyset 21.30, b \emptyset 8.54, c \emptyset 1.51 ...

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Thermodynamics began with
the study of heat and work
effects and relations
between heat and work. Some
early thermodynamics
problems were for very
practical problems. For
example, in a steam engine
heat is supplied to water to
create steam.

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**INTRODUCTION TO THE
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2.8* One mole of a monatomic
ideal gas at standard
temperature and pressure
(STP) undergoes the
following three processes:
1. at constant Pressure the
Temperature is doubled. 2.

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1.1* The plot of $V = V(P, T)$ for a gas is shown in
Fig. 1.1. Determine the
expressions of the two
second derivatives of the
volume of this plot.

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1.1* The plot of $V = V(P, T)$ for a gas is shown in
Fig. 1.1.

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Thus, the thermodynamic state of a system is uniquely fixed when a small number of macroscopic, independent variables are fixed. For example, consider a gas or a liquid of constant composition such as a pure gas or liquid. The three key variables are pressure, P , temperature, T , and volume, V .

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thermodynamics has much
wider applicability. In
material science, one is
normally not that interested

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in heat and work, but interested more the state of matter and how things might change when mixed, heated, pressurized, etc.

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