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This list gives you some of the most common conversion factors you need in thermodynamics. Acceleration: $1 \text{ m/s}^2 = 100 \text{ cm/s}^2$. Area: $1 \text{ m}^2 = 10^4 \text{ cm}^2 = 10^6 \text{ mm}^2$. Density: $1 \text{ g/cm}^3 = 1 \text{ kg/L} = 1,000 \text{ kg/m}^3$. Energy, heat, work, internal energy, enthalpy: $1 \text{ kJ} = 1,000 \text{ J} = 1,000 \text{ N}\cdot\text{m} = 1 \text{ kPa}\cdot\text{m}^3$. $1 \text{ kJ/kg} = 1,000 \text{ m}^2/\text{s}^2$.

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Basic Thermodynamic Formulas (Exam Equation Sheet) Control Mass (no mass flow across system boundaries) Conservation of mass: $\dot{m}_{in} = \dot{m}_{out}$. Conservation of energy (1st Law):

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$$p_1 - p_2 = \Delta p = \Delta p_1 + \Delta p_2 + \Delta p_3 = p_1 \Delta p_1 + p_2 \Delta p_2 - p_3 \Delta p_3$$

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Thermodynamics key facts (7/9) • Ideal gas

law • 1. st. form : $pV = nRT$. p = Pressure, V = Volume, n = number of

molecules, R = Boltzmann's constant,

T = temperature [in K] • 2. nd. form : $p = \frac{1}{3} n m \overline{v^2}$

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Thermodynamics key facts (3/9) • Heat energy

needed to change phase • The latent heat

L determines the energy needed to change

the phase of a mass m . • Units of L will be

J/kg . -1. - can be fusion or vaporization •

This energy is either absorbed (solid → liquid

→ gas) or released (gas → liquid → solid) • A

phase change takes place at constant

temperature.

~~Revision : Thermodynamics~~

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Lucky. My professor didn't let us use anything in our test other than a pencil. Calculator and the thermodynamic property table book. No cheat sheet, reference sheet. Nothing.

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Thermodynamics $\dot{Q} = \frac{\Delta T}{R}$ $\dot{P} = A v^3$ = rate of heat transfer $Q = thermal\ energy$ $A = Area\ of\ thermal\ conductivity$ $U = coefficient\ of\ heat\ conductivity\ (U-factor)$ $T = change\ in\ temperature$ $R = resistance\ to\ heat\ flow\ (R-value)$ $\dot{v} = velocity\ net = net\ power\ radiated = 5.6696 \times 10^{-8} \dot{T}_1, \dot{T}_2 = temperature\ at\ time\ 1, time\ 2$ $v = flow\ velocity$

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IV. The First Law of Thermodynamics as a Rate Equation A. Most general form $\frac{dE_{c.v.}}{dt} = \dot{Q}_{c.v.} - \dot{W}_{c.v.} + \dot{m}_{in} e_{in} - \dot{m}_{out} e_{out}$ rate of change of energy in c.v. = rate of heat added to c.v. rate of work done + rate of energy flow in to c.v. rate of energy flow

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out of c.v. B. For a steady flow process $d dt = 0$ and $m \cdot in = m \cdot out = m \cdot Q \cdot c.v. W \cdot$

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~~Thermodynamics: the study of energy, energy~~

~~transformations and its relation to matter.~~

~~The anal-ysis of thermal systems is achieved~~

~~through the application of the governing~~

~~conservation equations, namely Conservation~~

~~of Mass, Conservation of Energy (1st law of~~

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thermodynamics), the 2nd law of thermodynamics and the property relations.

~~Basic Concepts of Thermodynamics~~

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engineering work, pressures are often measured with respect to atmospheric pressure rather than with respect to absolute vacuum.
 $P_{\text{abs}} = P_{\text{atm}} + P_{\text{gauge}}$ In SI units the derived unit for pressure is the Pascal (Pa), where $1 \text{ Pa} = 1 \text{ N/m}^2$. This is very small for engineering purposes, so usually pressures are given in terms of kilopascals ($1 \text{ kPa} = 1000 \text{ Pa}$),

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